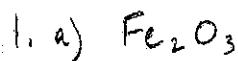
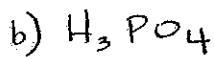


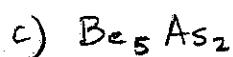
Mole Problems #0 - Answer Key



$$\left. \begin{array}{l} 2 \times \text{Fe} = 2 \text{ mol} \times 55.8 \text{ g/mol} = 111.6 \text{ g} \\ 3 \times \text{O} = 3 \text{ mol} \times 16.0 \text{ g/mol} = 48.0 \text{ g} \end{array} \right\} \text{ADD} = \boxed{159.6 \text{ g}}$$



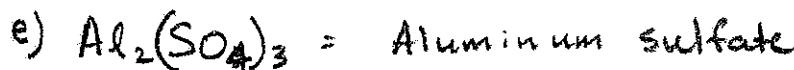
$$\left. \begin{array}{l} 3 \times \text{H} = 3 \text{ mol} \times 1.0 \text{ g/mol} = 3.0 \text{ g} \\ 1 \times \text{P} = 1 \text{ mol} \times 31.0 \text{ g/mol} = 31.0 \text{ g} \\ 4 \times \text{O} = 4 \text{ mol} \times 16.0 \text{ g/mol} = 64.0 \text{ g} \end{array} \right\} \text{ADD} = \boxed{98.0 \text{ g}}$$



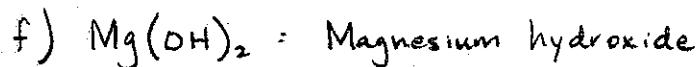
$$\left. \begin{array}{l} 5 \times \text{Be} = 5 \text{ mol} \times 9.0 \text{ g/mol} = 45.0 \text{ g} \\ 2 \times \text{As} = 2 \text{ mol} \times 74.9 \text{ g/mol} = 149.8 \text{ g} \end{array} \right\} \text{ADD} = \boxed{194.8 \text{ g}}$$



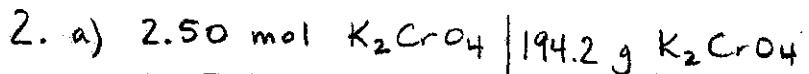
$$\left. \begin{array}{l} 2 \times \text{Rb} = 2 \text{ mol} \times 85.5 \text{ g/mol} = 171.0 \text{ g} \\ 1 \times \text{S} = 1 \text{ mol} \times 32.1 \text{ g/mol} = 32.1 \text{ g} \\ 3 \times \text{O} = 3 \text{ mol} \times 16.0 \text{ g/mol} = 48.0 \text{ g} \end{array} \right\} \text{ADD} = \boxed{251.1 \text{ g}}$$



$$\left. \begin{array}{l} 2 \times \text{Al} = 2 \text{ mol} \times 27.0 \text{ g/mol} = 54.0 \text{ g} \\ 3 \times \text{S} = 3 \text{ mol} \times 32.1 \text{ g/mol} = 96.3 \text{ g} \\ 12 \times \text{O} = 12 \text{ mol} \times 16.0 \text{ g/mol} = 192.0 \text{ g} \end{array} \right\} \text{ADD} = \boxed{342.3 \text{ g}}$$

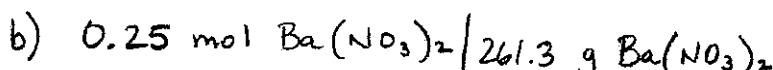


$$\left. \begin{array}{l} 2 \times \text{Mg} = 1 \text{ mol} \times 24.3 \text{ g/mol} = 24.3 \text{ g} \\ 2 \times \text{O} = 2 \text{ mol} \times 16.0 \text{ g/mol} = 32.0 \text{ g} \\ 2 \times \text{H} = 2 \text{ mol} \times 1.0 \text{ g/mol} = 2.0 \text{ g} \end{array} \right\} \text{ADD} = \boxed{58.3 \text{ g}}$$



$$\left| \begin{array}{c} 194.2 \text{ g } \text{K}_2\text{CrO}_4 \\ \hline 1 \text{ mol } \text{K}_2\text{CrO}_4 \end{array} \right| = 485.5 \text{ g } \text{K}_2\text{CrO}_4$$

$$= 486 \text{ g } \text{K}_2\text{CrO}_4 \text{ (SIG FIGS)}$$



$$\left| \begin{array}{c} 261.3 \text{ g } \text{Ba}(\text{NO}_3)_2 \\ \hline 1 \text{ mol } \text{Ba}(\text{NO}_3)_2 \end{array} \right| = 65.3 \text{ g } \text{Ba}(\text{NO}_3)_2$$

$$= 65 \text{ g } \text{Ba}(\text{NO}_3)_2$$

